

Numerical Analysis of the Self-Heating Behaviour of Coal Dust Accumulations



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Self-heating behaviour of dust accumulations is a multi-physics field coupled heat and mass transfer in the porous media. A typical experimental apparatus with a hot storage oven and mesh wire baskets has been taken as the study object. The influence of gas flow velocity, oxygen concentration and ambient temperature on the self-heating behaviour of the dry coal dust sample has been investigated with COMSOL Multiphysics®.

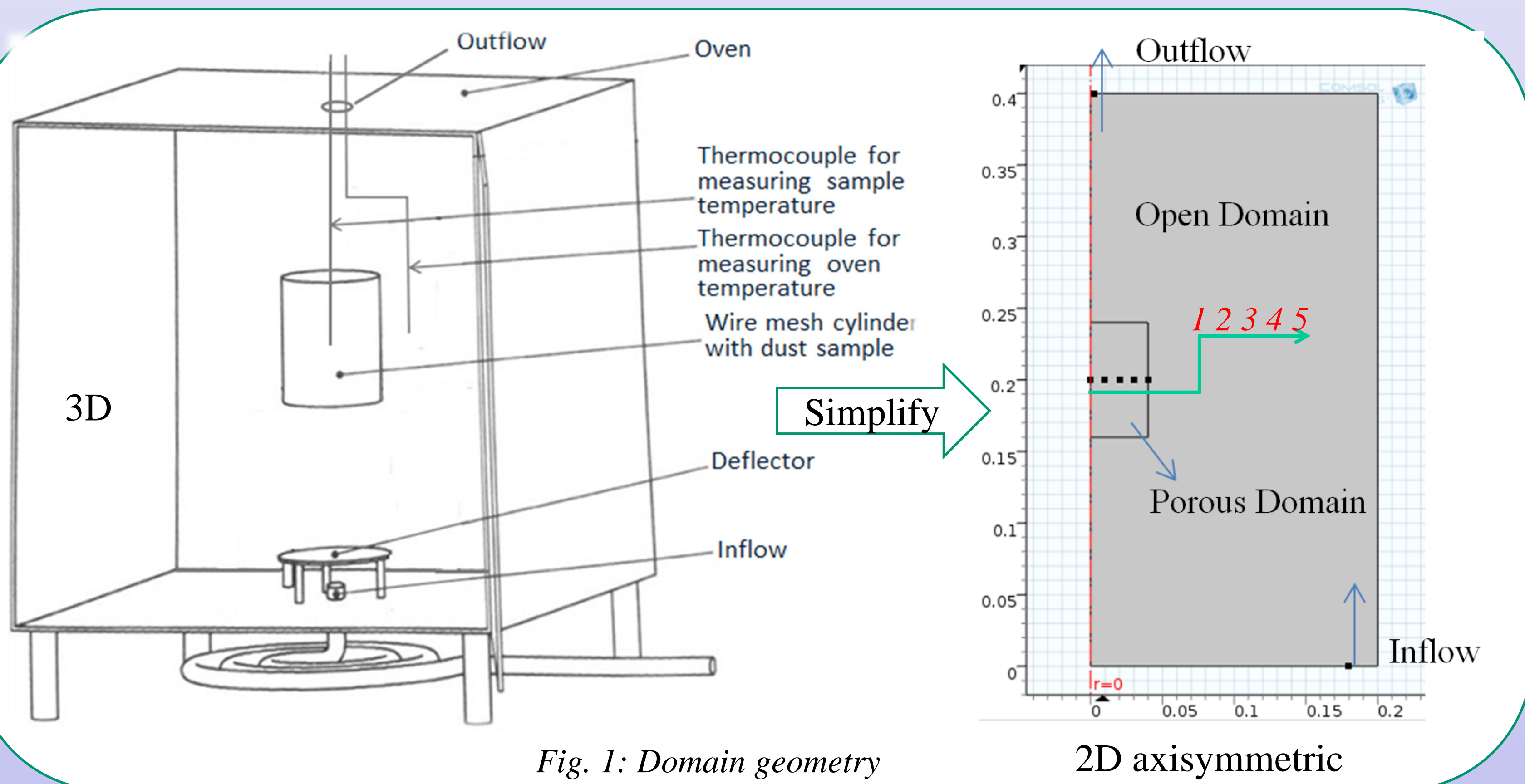


Fig. 1: Domain geometry

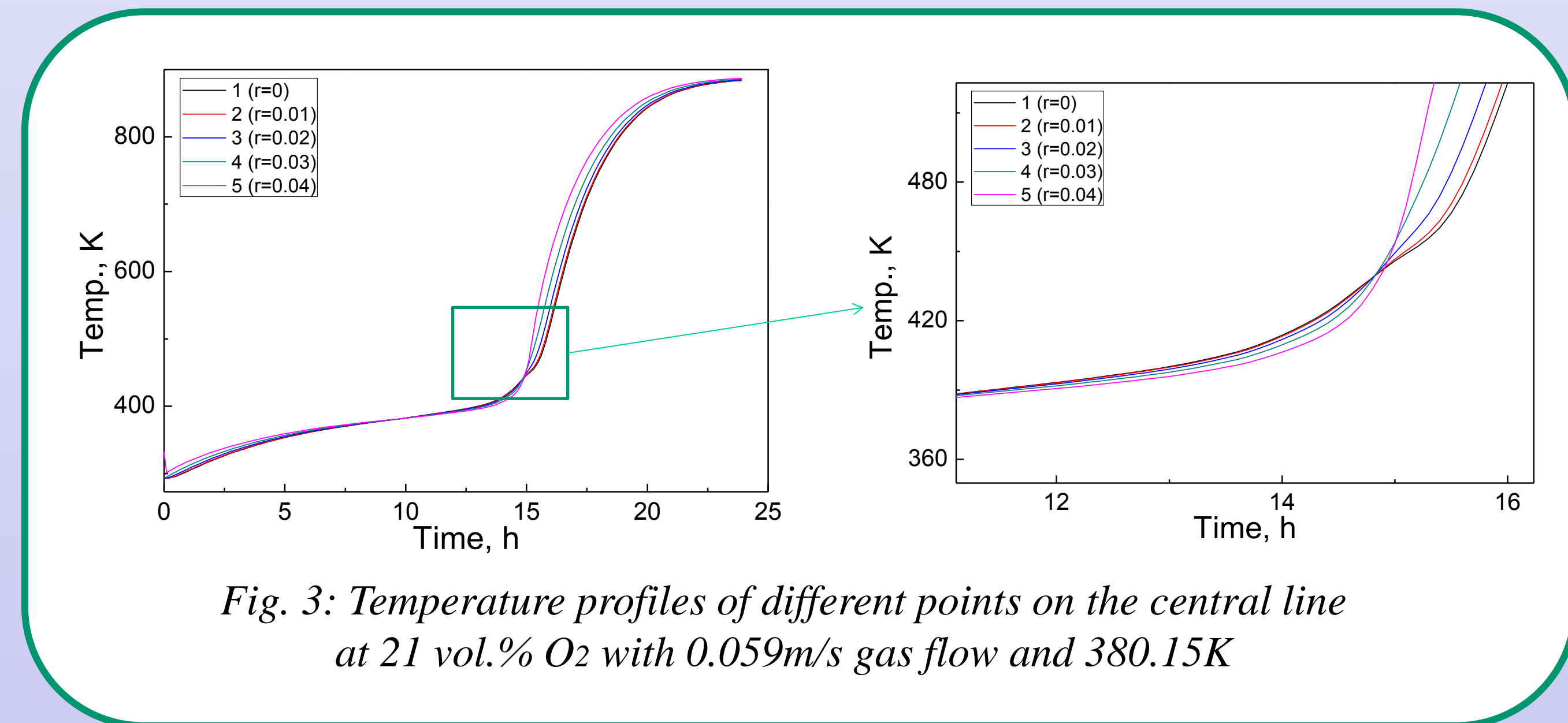


Fig. 3: Temperature profiles of different points on the central line at 21 vol.% O₂ with 0.059m/s gas flow and 380.15K

Chemical



$$\text{rate} = k \cdot c\text{O}_2 \quad k = A \exp\left(-\frac{E_A}{RT}\right)$$

Free and Porous Media Flow

$$\text{Open domain: } \rho \frac{\partial \mathbf{u}}{\partial t} = \nabla \cdot [-p\mathbf{I} + \mu(\nabla \mathbf{u} + (\nabla \mathbf{u})^T)]$$

Porous domain:

$$\frac{\rho}{\varepsilon_b} \frac{\partial \mathbf{u}}{\partial t} = \nabla \cdot \left[\frac{\mu}{\varepsilon_b} (\nabla \mathbf{u} + (\nabla \mathbf{u})^T) - p\mathbf{I} \right] - \frac{\mu}{\kappa} \mathbf{u}$$

$$\nabla \cdot (\rho \mathbf{u}) = 0$$

Species Transport in Porous Media

$$\text{Open domain: } \frac{\partial c_i}{\partial t} - \nabla \cdot (D_i \nabla c_i) + \mathbf{u} \cdot \nabla c_i = 0$$

Porous domain:

$$\varepsilon_b \frac{\partial c_i}{\partial t} + \nabla \cdot (-\varepsilon_b \tau_{F,i} D_{F,i} \nabla c_i) + \mathbf{u} \cdot \nabla c_i = R_i$$

Heat Transfer in Porous Media

$$\text{Open domain: } \rho C_p \frac{\partial T}{\partial t} + \rho C_p \mathbf{u} \cdot \nabla T = \nabla \cdot (k_g \nabla T) + Q$$

Porous domain:

$$(\rho C_p)_{\text{eq}} \frac{\partial T}{\partial t} + \rho C_p \mathbf{u} \cdot \nabla T = \nabla \cdot (k_{\text{eq}} \nabla T) + Q$$

$$(\rho C_p)_{\text{eq}} = (1 - \varepsilon_b) \rho_p C_{p,p} + \varepsilon_b \rho_g C_{p,g}$$

$$k_{\text{eq}} = (1 - \varepsilon_b) k_b + \varepsilon_b k_g$$

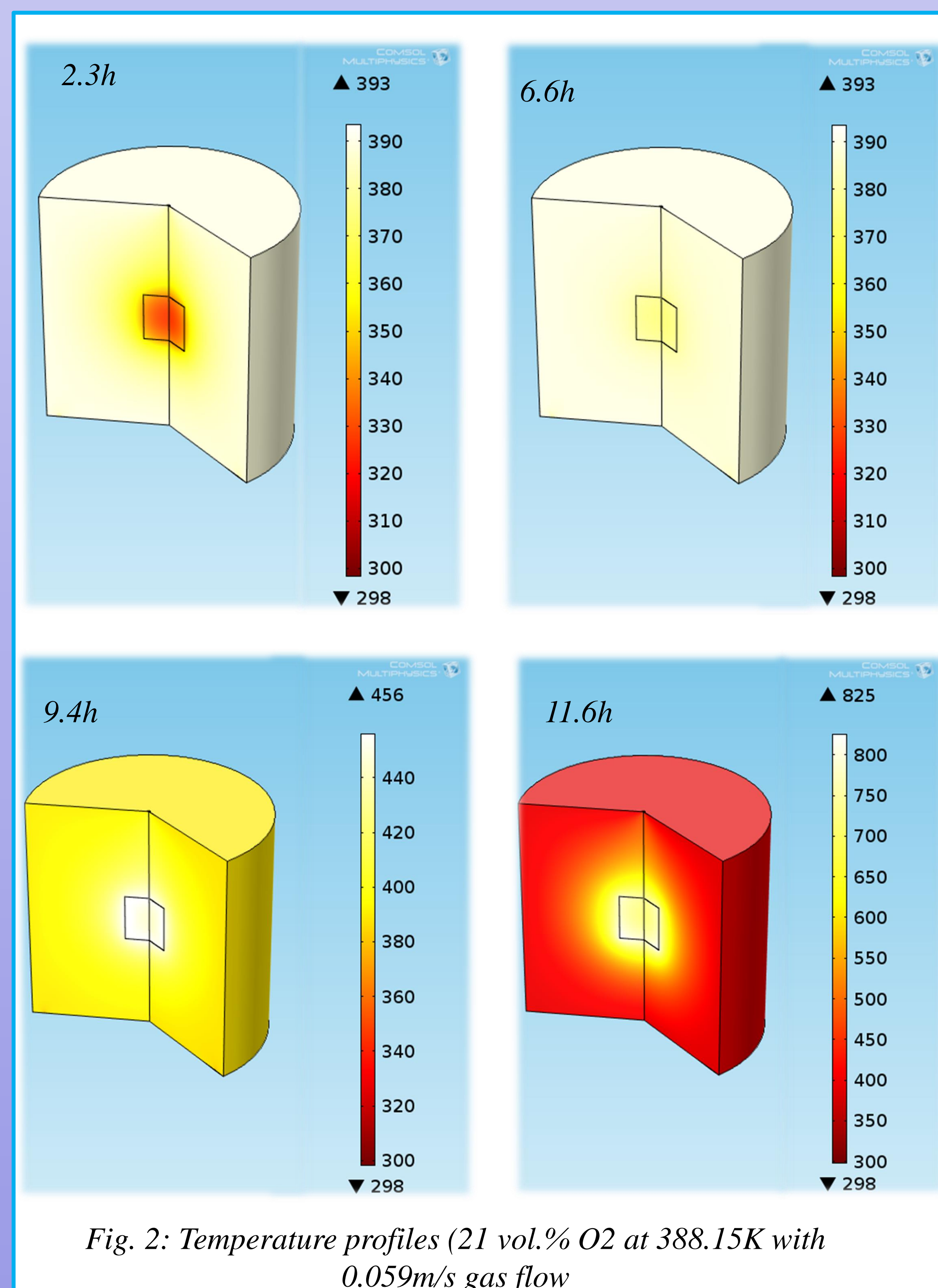


Fig. 2: Temperature profiles (21 vol.% O₂ at 388.15K with 0.059m/s gas flow)

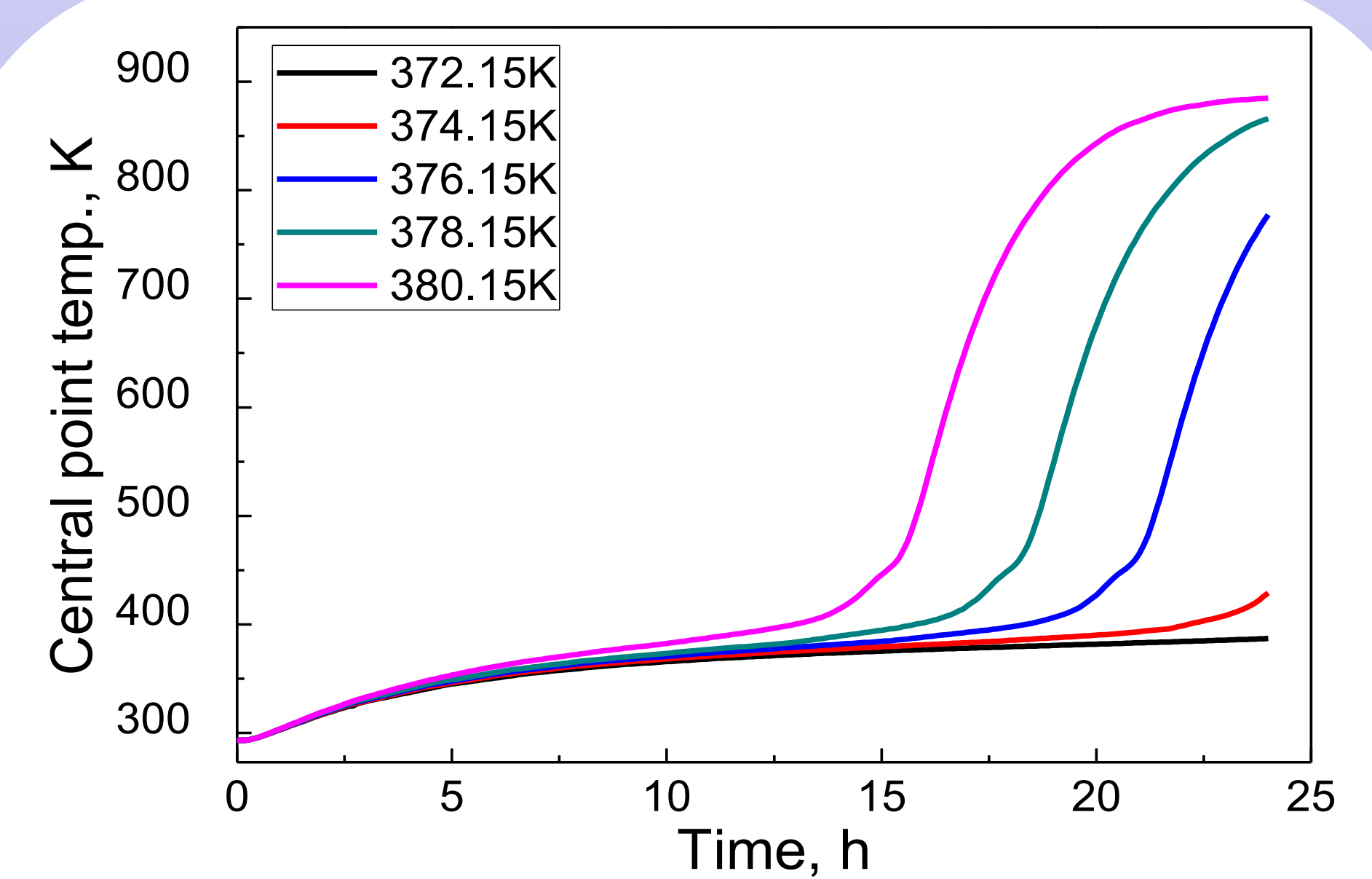


Fig. 4: Temperature evolution of the central point at 21 vol.% O₂ with 0.059m/s gas flow at various ambient temperatures

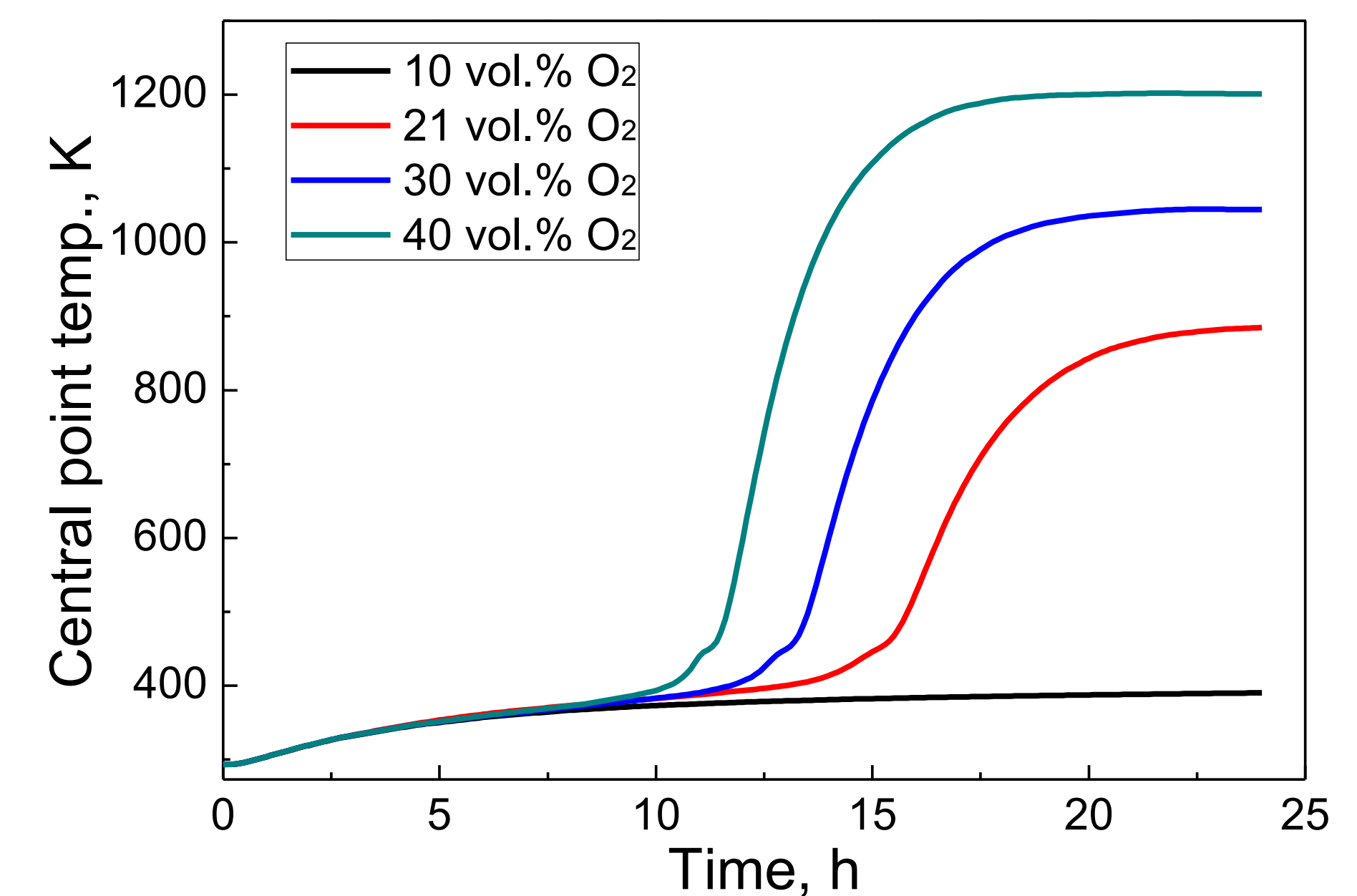


Fig. 5: Temperature evolution of the central point at 380.15K with 0.059m/s gas flow at various oxygen concentrations

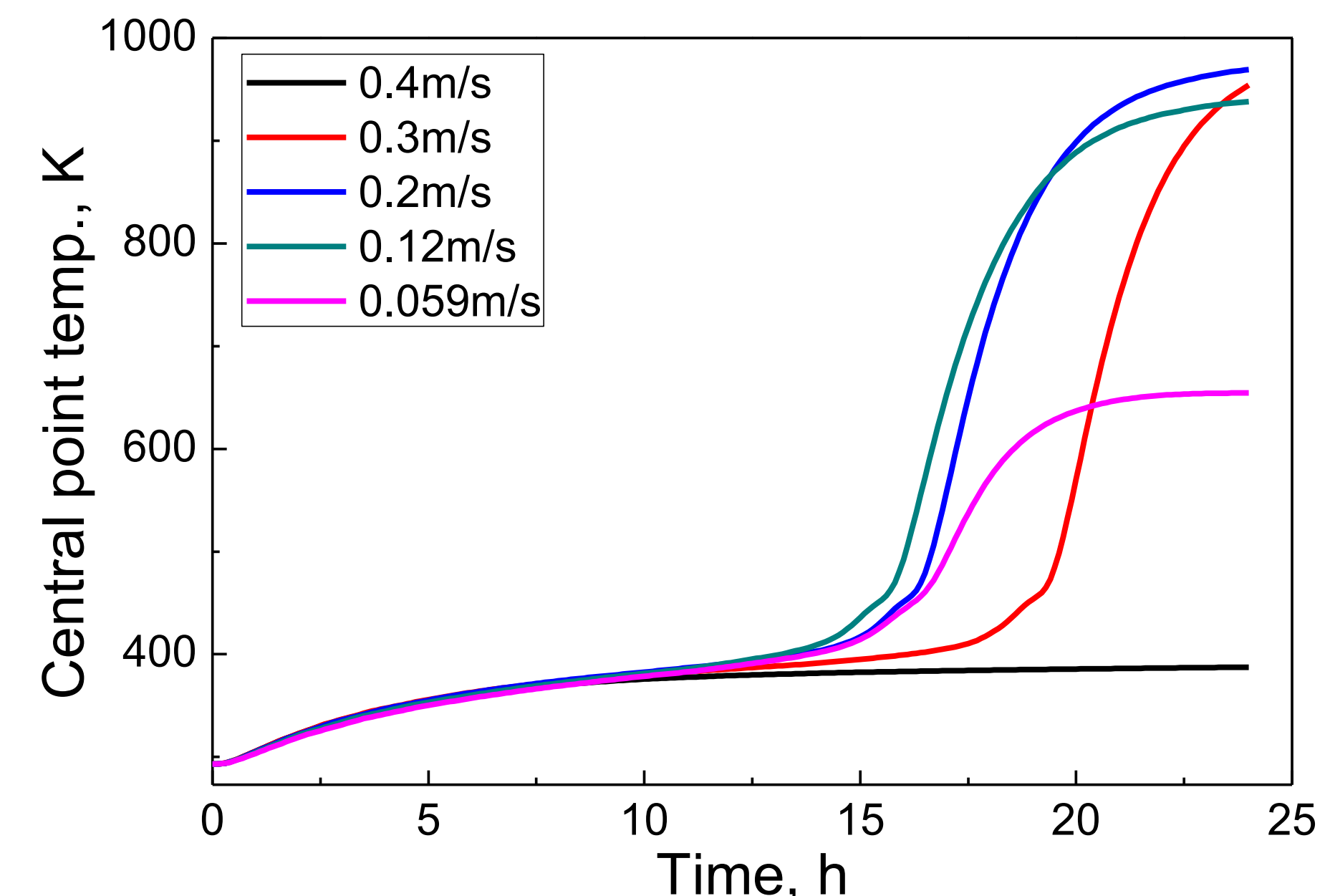


Fig. 6: Temperature evolution of the central point at 380.15K with 21 vol.% O₂ at various inflow velocities

Table 1: Input parameters

λ Thermal conductivity ($\text{W} \cdot \text{m}^{-1} \cdot \text{K}^{-1}$)	0.1
C_p Specific heat capacity ($\text{J} \cdot \text{kg}^{-1} \cdot \text{K}^{-1}$)	1.1e3
h Heat transfer coefficient ($\text{W} \cdot \text{m}^{-2} \cdot \text{K}^{-1}$)	
ρ_p Density of coal ($\text{kg} \cdot \text{m}^{-3}$)	1.2e3
ρ_b Bulk Density ($\text{kg} \cdot \text{m}^{-3}$)	600
$\varepsilon = \frac{\rho_p - \rho_b}{\rho_p - \rho_g} \approx 1 - \frac{\rho_b}{\rho_p}$ Porosity	0.5
E_a Activation energy in air oxidation ($\text{kJ} \cdot \text{mol}^{-1}$)	98
A pre-exponential factor (s^{-1})	1e7
M_{fuel} molar weight of coal ($\text{g} \cdot \text{mol}^{-1}$)	4e3
ΔH Heat of reaction ($\text{kJ} \cdot \text{mol}^{-1} \text{O}_2$)	400
r_p average coal particle radius (m)	1e-3

Conclusions and future work

- ✓ Ignition delay time decreases with increasing ambient temperature and oxygen concentration;
- ✓ The maximum temperature can be reached after ignition increases with increasing ambient oxygen concentration;
- ✓ The inflow velocity has both positive and negative effects on self-heating. The critical velocity is between 0.12 to 2m/s.
- ❖ The influence of particle size and depletion of coal should be taken into consideration.

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References

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